

Shapes of molecules (Level 2) examiners tips: Read these please!

- use molypods, plastercine & toothpicks or balloons etc to help you to visualise the shapes in 3D
- make and rotate interactive animations of molecules
- use [molecule3d](#) (*see link below*) to reinforce your visual images of the different shapes
- recognise and state the number of regions (areas) or electron repulsion around a central atom
- learn the names of the shapes and their angles
- learn them, no excuses!
- you are determining the shape around the central atom
- yes, an alternative accepted word for trigonal is triangular
- see below a summary of the shapes of molecules for Level 2, also available to download as a Powerpoint

Also..."don't be daft", what NOT to do

a single or double or triple bond are all ONE REGION of negative charge
the trigonal planar shape is VERY, VERY different to trigonal pyramidal shape